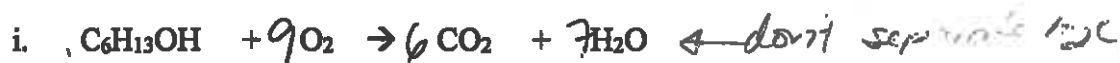
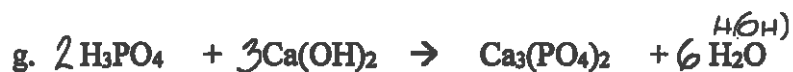


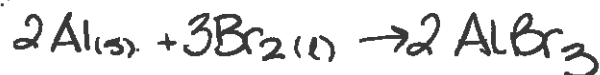
Key

**Balancing Chemical Equations**

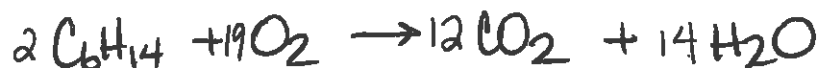
1. Balance the following equations (1 mark each = 10 marks)

2. Write a balanced chemical equation for each of the following. Don't forget *diatomic* elements!

a. aluminum metal reacts with bromine to form aluminum bromide



b. hydrochloric acid neutralizes aluminum hydroxide to form water &amp; aluminum chloride

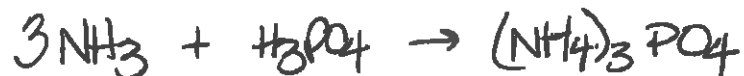
c. hexane ( $\text{C}_6\text{H}_{14}$ ) burns in oxygen to produce carbon dioxide and waterd. carbon dioxide and water are reacted to produce glucose ( $\text{C}_6\text{H}_{12}\text{O}_6$ ) and oxygen in photosynthesis

## Chemistry 11

- e. aluminum nitrate reacts with lithium sulphate to form aluminum sulphate and lithium nitrate



- f. ammonia and phosphoric acid react to form ammonium phosphate



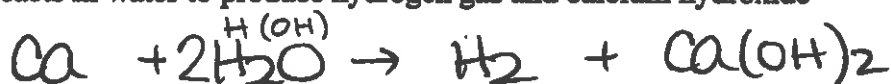
- g. nitrogen dioxide reacts with water to form nitric acid and nitrogen monoxide



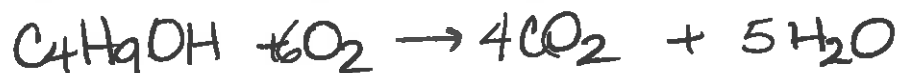
- h. bromine reacts with sodium iodide to produce iodine and sodium bromide



- i. calcium reacts in water to produce hydrogen gas and calcium hydroxide



- j. butanol ( $\text{C}_4\text{H}_9\text{OH}$ ) burns in oxygen to produce carbon dioxide and water



- k. hydrogen peroxide decomposes to form water and oxygen gas



- l. aluminum bromide reacts with ammonium dichromate to produce aluminum dichromate and ammonium bromide

