

Mixed NOMENCLATURE Practice

Molecular Compounds, Ionic Compounds, & Acids

NAME THE FOLLOWING COMPOUNDS:

- | | | | |
|---------------------------------------|-------|--------------------------------------|-------|
| 1. BaSO_3 | _____ | 14. AgNO_3 | _____ |
| 2. $(\text{NH}_4)_3\text{PO}_4$ | _____ | 15. As_2O_5 | _____ |
| 3. PBr_5 | _____ | 16. Fe_2O_3 | _____ |
| 4. MgSO_4 | _____ | 17. HClO | _____ |
| 5. CaO | _____ | 18. N_2O_3 | _____ |
| 6. H_3PO_4 | _____ | 19. HF | _____ |
| 7. $\text{Na}_2\text{Cr}_2\text{O}_7$ | _____ | 20. $\text{H}_2\text{C}_2\text{O}_4$ | _____ |
| 8. MgO | _____ | 21. NaHCO_3 | _____ |
| 9. SO_3 | _____ | 22. SiBr_4 | _____ |
| 10. $\text{Cu}(\text{NO}_3)_2$ | _____ | 23. CuCl_2 | _____ |
| 11. HI | _____ | 24. HNO_2 | _____ |
| 12. N_2O | _____ | 25. SnO_2 | _____ |
| 13. MnO | _____ | 26. BaCrO_4 | _____ |

WRITE FORMULAS FOR THE FOLLOWING COMPOUNDS:

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|-----------------------------|-------|----------------------------|-------|
| 27. hydrobromic acid | _____ | 40. diphosphorus pentoxide | _____ |
| 28. chromium(III) carbonate | _____ | 41. sulfurous acid | _____ |
| 29. magnesium sulfide | _____ | 42. lead(II) nitrate | _____ |
| 30. iodine trichloride | _____ | 43. dihydrogen monoxide | _____ |
| 31. lithium hydride | _____ | 44. sodium oxalate | _____ |
| 32. ammonium hydroxide | _____ | 45. perchloric acid | _____ |
| 33. calcium chloride | _____ | 46. chlorous acid | _____ |
| 34. hydroselenic acid | _____ | 47. silicon dioxide | _____ |
| 35. iron(II) nitride | _____ | 48. carbonic acid | _____ |
| 36. aluminum hydroxide | _____ | 49. sodium chlorate | _____ |
| 37. tin(II) fluoride | _____ | 50. xenon hexafluoride | _____ |
| 38. sulfur tetrachloride | _____ | 51. nickel (II) nitrate | _____ |
| 39. mercury(II) iodide | _____ | 52. potassium perchlorate | _____ |