

Name

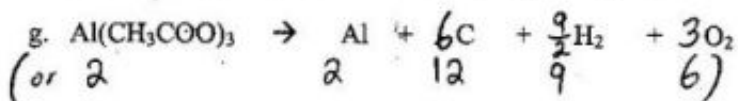
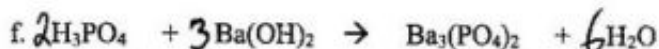
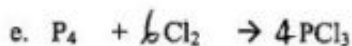
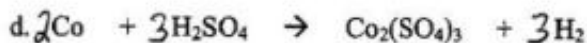
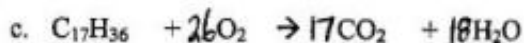
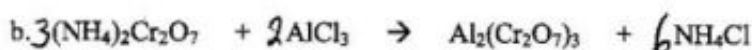
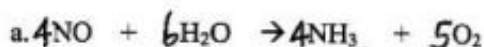
KEY

Date

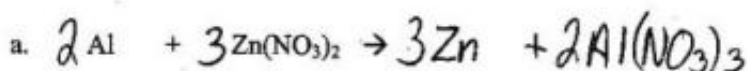
Chemistry 11

Review of Unit 6

1. Balance the following equations:

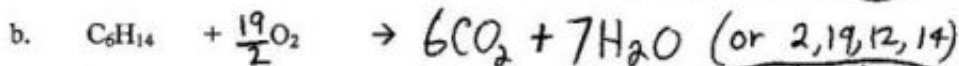


2. Complete, balance and classify the following equations as synthesis, decomposition, single replacement, double replacement, neutralization or combustion.



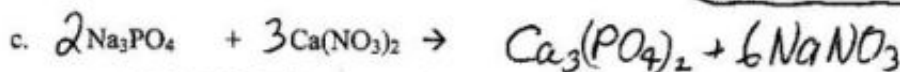
Reaction Type

SR



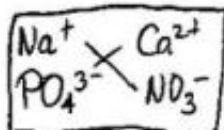
Reaction Type

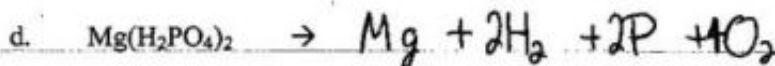
Combustion



Reaction Type

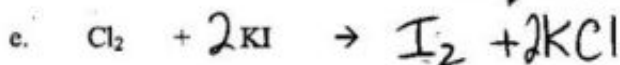
DR





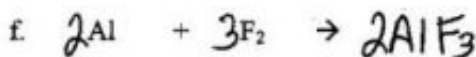
Reaction Type

decomposition



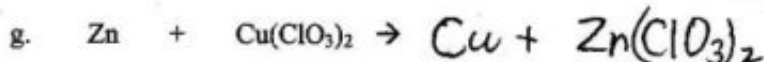
Reaction Type

SR



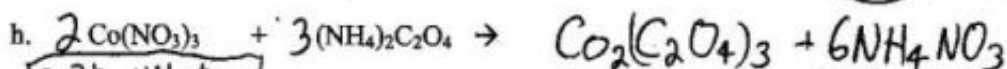
Reaction Type

synthesis



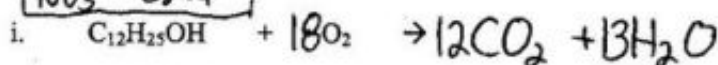
Reaction Type

SR



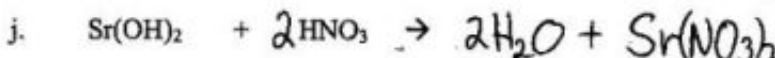
Reaction Type

DR



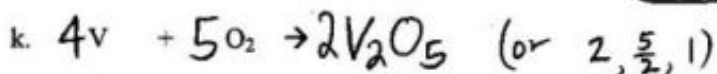
Reaction Type

Combustion



Reaction Type

neutralization or DR



(Assume combining capacity of V is 5+)

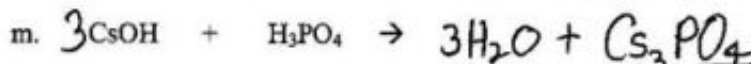
Reaction Type

Synthesis



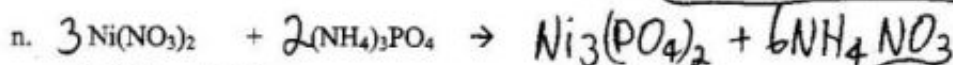
Reaction Type

decomposition



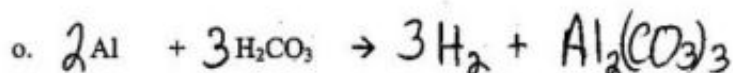
Reaction Type

neutralization or DR



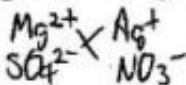
Reaction Type

DR



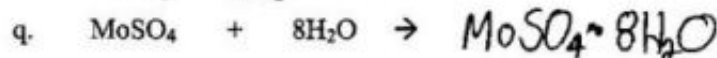
Reaction Type

SR



Reaction Type

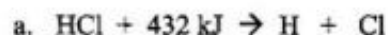
DR



Reaction Type

synthesis (hydration)

3. State whether each of the following are *exothermic* or *endothermic*.



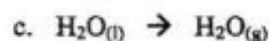
Answer

endo



Answer

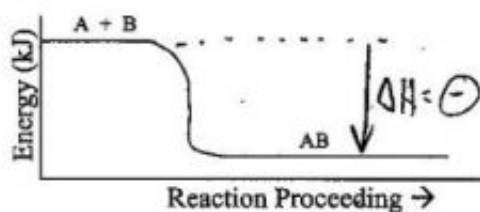
exo



Answer

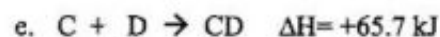
endo

d.



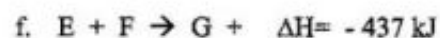
Answer

exo



Answer

endo



Answer

exo

4. In an *exothermic* reaction, the surroundings get (warmer/cooler)

warmer

5. Define *enthalpy*

stored chemical energy in a substance