Chem 11 **Matter Review Package**

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Hebden: Unit III

**UNIT III: MATTER**

*In addition to these questions, make sure to look at the definitions and examples in your notes.*

1. A mixture (**is / is not**) composed of two or more substances.
2. **True or False**: An element can be broken down into a simpler substance.
3. From the following list, circle the ones that are elements:

silver
water
oxygen
air
carbon dioxide
hydrogen
gold
sulphur
alcohol
carbon
sugar
magnesium
chromium
nitrogen
salt
nickel

1. Draw the classification of matter flow chart. Be sure to include the following: **matter, suspensions, compounds, mixtures, pure substances, elements, solutions,** and **mechanical mixtures**. Write a characteristic below each word. Give examples of the two kinds of pure substances and the three kinds of mixtures.
2. Classify the following as pure substances or mixtures.

air \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
water \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
mercury \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
gasoline \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
sugar \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
oxygen \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
gold \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
salt water \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. Classify the following as heterogenous or as homogeneous (assume they are all mixtures).

salt water \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
aluminum foil \_\_\_\_\_\_\_\_\_\_\_\_\_\_
tap water \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
tossed salad \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
unfiltered air\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
an apple \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
iron with rust \_\_\_\_\_\_\_\_\_\_\_\_\_\_
wood \_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. a) Explain the principles behind how solvent extraction works.

 b) Explain the principles behind how distillation works.

1. Classify the following properties of matter as physical or chemical.

Colour \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
Density \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
Burns easily (flammable) \_\_\_\_\_\_\_\_\_\_\_
Boils at 450°C \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
Melts at 145°C \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Dissolves in water \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. Classify the following as an intensive property (**I**) or an extensive property (**E**).

Mass \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
Density \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
Melting Point \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
Colour \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
Volume \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Length \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. Answer the questions below.
2. Fill in the table below by checking the appropriate column.
3. Draw a heating curve for both a pure substance and mixture below (on separate graphs). Be sure to include the following:x and y axis titles

Label: solid, liquid and gas states & melting/freezing/recrystallization point & evaporation/condensation/ boiling points

**Heating Curve of a Pure Substance**